In "Detection of Gross Errors in Process Data" by R.S.H. Mah and A. C. Tamhane [*AIChE J.*, **28**, 828 (1982)],

line 33, left column, p. 828 should read "... the vector of true flow rates (μ), ..." not "... the vector of true flow rates s(μ), ...".

Line 32, left column, p. 829 should read "... $Z_{\beta/2} > Z_{\alpha/2}$" not "... $Z_{\beta/2} > Z_{\beta/2}$".

The Q's and δ in Eq. (19) should be printed in bold-face characters.

The first equation in the right column, p. 829 should be labeled Eq. (21).

In "Hydrogenolysis and Hydrogenation of Dibenzothiophene Catalyzed by Sulfided $CoO\text{-}MoO_3/\gamma\text{-}Al_2O_3$: The Reaction Kinetics," by D. H. Broderick and B. C. Gates, [*AIChE*, **27**, 663 (1981)] the following corrections should be made in Table 4, column 1:

"0.153 kg mol/m^3" should have been "0.0153 kg mol/m^3"

and

"0.617 kg mol/m^3" should have been "0.0617 kg mol/m^3."

In "Effects of London-van der Walls Forces on the Thinning of a Dimpled Liquid Film as a Small Drop or Bubble Approaches a Horizontal Solid Plane" by Jing-Den Chen and J. C. Slattery [*AIChE J.*, **28**, 955 (1982)] the following discussion section and table were added by the authors after the issue was closed:

DISCUSSION

In the systems studied by Platikanov (1964), a bubble was pressed against a solid surface through a continuous liquid phase and the draining liquid film that was formed evolved into a uniform film as equilibrium was approached. The existance of this uniform film was the result of a balance between the positive disjoining pressure and the capillary pressure (Sheludko, 1967; Buscall and Ottewill, 1975). Blake (1975) demonstrated that a positive disjoining pressure attributable only to London-van der Waals forces was sufficient to form such an equilibrium film.

Using values of B^* and the final film thickness h^*_∞ obtained from our computations, we find in Table 1 good comparisons between the capillary pressure and the London-van der Waals disjoining pressure for the three systems studied by Platikanov (1964).

TABLE 1. COMPARISONS OF CAPILLARY PRESSURE AND DISJOINING PRESSURE

system	$\frac{2\gamma_0^*}{R_b^*}$ (Pa)	$\frac{-B^*}{h_\infty^{*4}}$ (Pa)
0.1 N KCl solution-air	60.6	61.7
aniline-air	30.8	31.1
ethanol-air	18.6	18.6

In "Studies of Mold Filling and Curing in the Reaction Injection Molding Process" by J. M. Castro and C. W. Macosko [*AIChE J.*, **28**, 250 (1982)] the following adjustments to the notation section on page 260 are necessary:

$$k^* = \text{dimensionless reaction rate} = \frac{k}{k_{T_m}} \frac{C_g^*}{(1 - C_g^*)}$$

$$= \frac{C_g^*}{1 - C_g^*} e^{-E^*\Delta T_{ad}\left[\frac{1}{T} - \frac{1}{T_m}\right]}$$

$$\eta^* = \text{dimensionless viscosity} = \frac{\eta}{\eta_{T_0}}$$

$$= \left[\frac{C_g^*}{C_g^* - C^*}\right]^{A + BC^*} e^{-E_\eta \Delta T_{ad}\left[\frac{1}{T} - \frac{1}{T_0}\right]}$$

In "Drop Size and Continuous-Phase Mass Transfer in Agitated Vessels" by A. H. P. Skelland and Jai Moon Lee, [*AIChE J.*, **27**, 99 (1981):

1) Page 109, col. 2, 2nd paragraph, should read:

To obtain equal capacity coefficients in two geometrically similar units of different size, when $\phi_1 = \phi_2$ and the relevant physical properties have respectively the same values on the two scales, equating the righthand side of Eq. 51 for systems 1 and 2 leads to

$$\frac{N_1}{N_2} = \left(\frac{d_{I_2}}{d_{I_1}}\right)^{0.966} \tag{60}$$

2) Equation (61) should read

$$\frac{P_1}{P_2} = \left(\frac{d_{I_1}}{d_{I_2}}\right)^5 \left(\frac{N_1}{N_2}\right)^3 = \left(\frac{d_{I_1}}{d_{I_2}}\right)^{2.102} \tag{61}$$

3) Equation (62) in corrected form is

$$\frac{P_1/Vol_1}{P_2/Vol_2} = \frac{P_1}{P_2}\left(\frac{d_{I_2}}{d_{I_1}}\right)^3 = \left(\frac{d_{I_2}}{d_{I_1}}\right)^{0.898} \tag{62}$$

4) The final paragraph should then read:

Eq. 62 shows the power input per unit volume to make $(k_c a)_1 = (k_c a)_2$ decreases with increasing d_I for the range of d_I/T studied.